

$A = LiNC$



~~B = NO₂~~



~~w(X в B) = 20.2%~~



$w(Li) = \frac{7}{36} \cdot 100\% = 20.2\%$

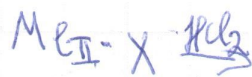


$m(B \cdot n H_2O) = 9.04$

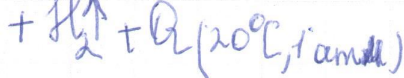


S1

$m[Me(II) \text{ и } X] = 272$



4811



$272 - 42 = 142 + O_2 \rightarrow 19.202$

$Me_{II} = Al = 102$

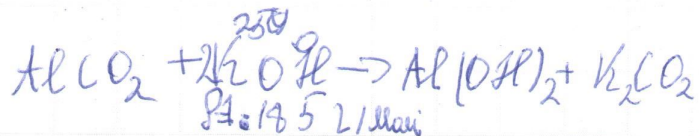
$w(Al) = \frac{10}{27} \cdot 100\% = 37\%$

$X = CO_2 = 172$

$\frac{12}{27} \cdot 272 = 122$
 $27 - x$

$w(CO_2) = \frac{17}{27} \cdot 100\% = 63\%$

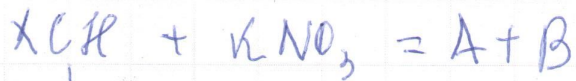
$x = 10 \quad 27 - 10 = 17$



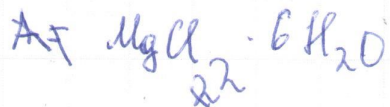
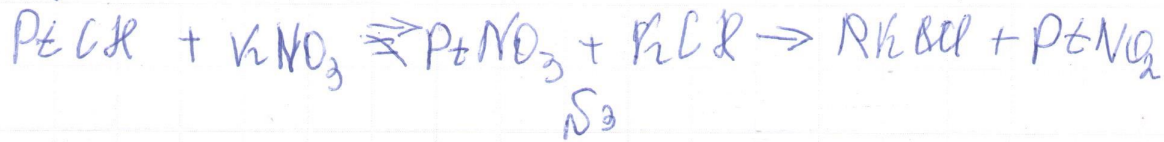
$v(AlCO_2) = \frac{m}{224}$

$v(AlCO_2) = \frac{27}{224} = 1.21$

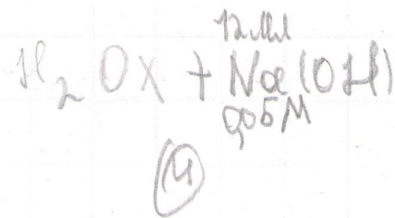
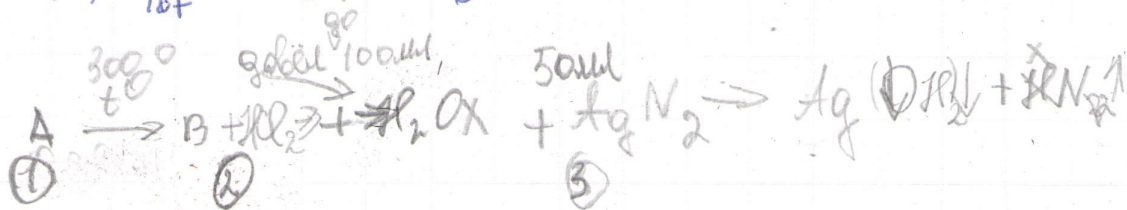
S4



↓



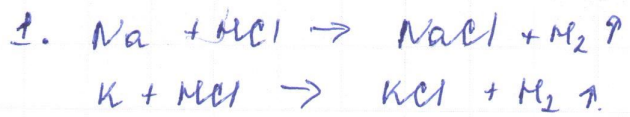
$$w(Cl) = \frac{35}{164} \cdot 100\% = 21\%$$



Задача 1.



$m(\text{Me(II)} \text{ и } \text{X}) = 27 \text{ г}$
 4.81 (V)



$$w = \frac{M(\text{?})}{M(\text{в-ва})} \cdot 100\%$$

$$w(\text{Na}) = \frac{22}{27} \cdot 100 = 81.5\%$$

$$w(\text{K}) = \frac{39}{27} \cdot 100 = 144.4\%$$

2. $w(\text{KOH}) = 25\%$

$$\rho(\text{KOH}) = 1.1852 \text{ г/мл}$$

$V = ?$

решение

$$[V = m \cdot \rho]$$

$$V = 56 \cdot 1.185 = 66.26 \text{ мл}$$

ответ: $V(\text{KOH}) = 66.26 \text{ мл}$

Задача 2.

- А - бинарное соединение
- Б - земной окислитель газ.
- В - безводная бинарная соль
- Р - токсичный газ.

$$w(\text{вчх}) = 20.2\%$$